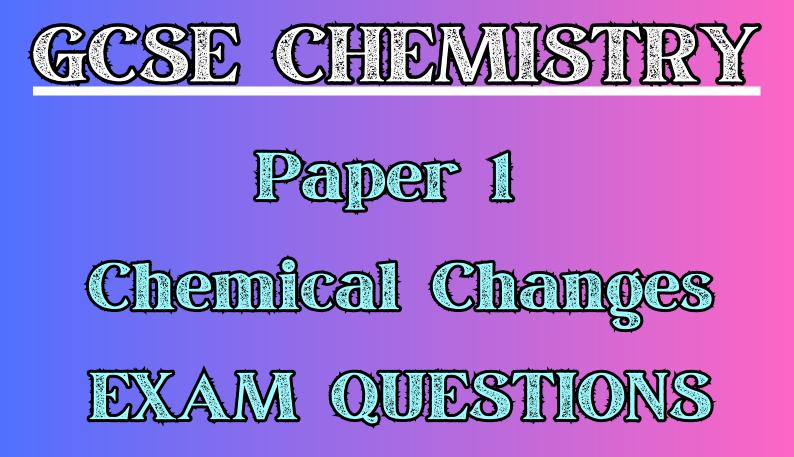
Grade Nine Guides



Exam questions per topic

Aiming at grade 8/9



Answers in No Waffle GCSE videos :)

Chemical changes Required practical 1: making salts	 Write an ionic equation for the neutralisation of an acid and an alkali [2] Required practical: how to prepare a salt from insoluble metal carbonate/ oxide an acid; [A] Describe a method for making pure crystals of copper sulphate from copper carbonate and dilute sulfuric acid. 	
Required practical 2:neutralisation [titration]	 Required practical: how to prepare a salt from soluble metal carbonate/ oxide and acid B] Describe a method for making crystals of potassium chloride from potassium hydroxide and hydrochloric acid [titration] 	
Required practical 3: electrolysis	 C: Explain why the student should use a pipette to measure the dilute sulfuric acid and a box burette to measure the sodium hydroxide solution. [2 marks] Required practical to: electrolysis [6] 	

Answers in No Waffle GCSE videos :)

 Explain what happens at each electrolysis of aluminium oxide Explain why a mixture is used as the electrolyte instead of using only aluminium oxide. Explain why the positive electrode must be continually replaced. What are the differences between strong and weak acids What is the differences between nor low concentrations of acids Explain why the plot of an acid depends on: the strength of the acid, the concentration of the acid the concentration of the acid the concentration of the acid Dilute hydrochloric acid is a strong acid. Explain why an acid can be described as both strong and dilute. What is the difference between a chemical cell and an electrolyte is the acid and electrolyte is the strong and dilute. What is the difference between a chemical cell and an electrolyte is the strong and dilute. Explain why the excess hydrogen 	[2] [4]	 Suggest how in electrolysis the students could find the total collected Compare the use of nickel and platinum for electrodes in plugs 	
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chemical cell and an electrolyte [2]		acid. Explain why an acid can be described as both strong and dilute.	
Explain why the excess hydrogen	[2]	chemical cell and an electrolyte	
	^	Explain why the excess hydrogen	

must be burned off

[1]